

CHEM 1515.002
Exam II
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March 4, 1998

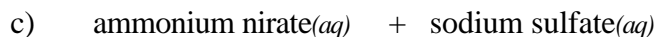
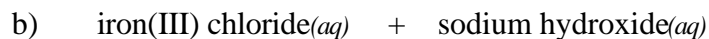
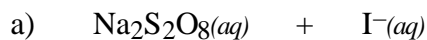
Name _____
TA's Name _____
Lab Section _____

INSTRUCTIONS:

1. This examination consists of a total of 8 different pages. The last three pages include a periodic table, a table of vapor pressures for water, a solubility table, a table of ionic structures and packing and an activity series. All work should be done in this booklet.
2. PRINT your name, TA's name and your lab section number now in the space at the top of this sheet. DO NOT SEPARATE THESE PAGES.
3. Answer all questions that you can and whenever called for show your work clearly. Your method of solving problems should pattern the approach used in lecture. You do not have to show your work for the multiple choice or short answer questions.
4. No credit will be awarded if your work is not shown in problems 3, 4, 5, 6 and 7.
5. Point values are shown next to the problem number.
6. Budget your time for each of the questions. Some problems may have a low point value yet be very challenging. If you do not recognize the solution to a question quickly, skip it, and return to the question after completing the easier problems.
7. Look through the exam before beginning; plan your work; then begin.
8. **Relax** and do well.

	Page 2	Page 3	Page 4	Page 5	TOTAL
SCORES	<u>(24)</u>	<u>(26)</u>	<u>(34)</u>	<u>(16)</u>	<u>(100)</u>

(12) 1. Write the chemical formula(s) of the product(s) and balance the following reactions. Identify all products phases as either (g)as, (l)iquid, (s)olid or (aq)ueous. Soluble ionic compounds should be written in the form of their component ions.



(6) 2. Write the ionic and net ionic chemical equations for 1a) and 1b).

1a)

Ionic equation:

Net Ionic equation:

1b)

Ionic equation:

Net Ionic equation:

(6) 3. Describe how you would prepare 250.00 mLs of an aqueous solution of NaOH which is 0.200 M. (Include any safety information which would be important to the person preparing the solution.)

- (26) 4. A solution is prepared containing 8.00 g of trichloroacetic acid, CCl_3COOH , in 92.0 g of water. This solution is also 0.509 M.
- (6) a) calculate the mol fraction of trichloroacetic acid in the solution.
- (6) b) calculate the freezing point of the solution (assuming the CCl_3COOH does not dissociate into ions.)
- (4) c) the experimental freezing point of this solution is found to be -1.69°C . Is trichloroacetic acid a strong, weak or nonelectrolyte in water? Explain.
- (6) d) Identify the chemical specie(s) in an aqueous solution of trichloroacetic acid. (NOTE: you need not include water, just the specie(s) associated with trichloroacetic acid.) Which is/are present in the highest concentration? Explain.
- (4) e) Calculate the density of this solution.

(34) 5. The 2nd order rate constant for the gas phase decomposition



of nitrosyl chloride, NOCl, is $1.18 \times 10^{-2} \text{ M}^{-1}\cdot\text{s}^{-1}$ at 450 K.

(8) a) If the initial concentration of NOCl is 0.500 M calculate the concentration of NOCl after 56.5 seconds.

(4) b) What is the fraction of the original NOCl is remaining?

(6) c) How long will it take for the concentration of NOCl to fall to one-half its initial value (as given in a)?

(8) d) If the activation energy, E_a , for the decomposition reaction is $102 \text{ kJ}\cdot\text{mol}^{-1}$, calculate the rate constant for the reaction at 600 K.

(4) e) Suggest a reaction mechanism for the decomposition reaction.

(4) f) Draw the activated complex for the slow step in your reaction mechanism (be sure to label where bond breaking and forming are occurring).

- (8) 6. Camphor, $C_{10}H_{16}O$, is frequently used to determine the molar mass of organic compounds using freezing point depression data because of its large freezing point constant, $37.7 \text{ m} \cdot ^\circ\text{C}^{-1}$. An unknown organic compound of mass 0.840 g lowered the freezing point of 25.0 g of camphor by $7.6 \text{ }^\circ\text{C}$. What is the molar mass of the organic compound?

- (8) 7. The reaction



was studied at $760 \text{ }^\circ\text{C}$. 0.200 mol of PCl_5 are placed in a 1.00 liter container and allowed to decompose. After equilibrium was established the concentration of PCl_3 was found to be 0.195 M. Calculate the equilibrium constant for the reaction at this temperature.

Periodic Table of the Elements

	IA																	VIII A	
1	1 H 1.008																		2 He 4.00
2	3 Li 6.94	4 Be 9.01										5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 19.00	10 Ne 20.18		
3	11 Na 22.99	12 Mg 24.30										13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.06	17 Cl 35.45	18 Ar 39.95		
4	19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.38	31 Ga 69.72	32 Ge 72.59	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80	
5	37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc (98)	44 Ru 101.1	45 Rh 102.9	46 Pd 106.4	47 Ag 107.9	48 Cd 112.4	49 In 114.8	50 Sn 118.7	51 Sb 121.8	52 Te 127.6	53 I 126.9	54 Xe 131.3	
6	55 Cs 132.9	56 Ba 137.3	57 La 138.9	72 Hf 178.5	73 Ta 180.9	74 W 183.8	75 Re 186.2	76 Os 190.2	77 Ir 192.2	78 Pt 195.1	79 Au 197.0	80 Hg 200.6	81 Tl 204.4	82 Pb 207.2	83 Bi 209.0	84 Po (209)	85 At (210)	86 Rn (222)	
7	87 Fr (223)	88 Ra 226.0	89 Ac 227.0	104 Rf (261)	105 Db (262)	106 Sg (263)	107 Bh (262)	108 Hs (265)	109 Mt (266)										

Lanthanides	58 Ce 140.1	59 Pr 140.9	60 Nd 144.2	61 Pm (145)	62 Sm 150.4	63 Eu 152.0	64 Gd 157.2	65 Tb 158.9	66 Dy 162.5	67 Ho 164.9	68 Er 167.3	69 Tm 168.9	70 Yb 173.0	71 Lu 175.0
Actinides	90 Th 232.0	91 Pa 231.0	92 U 238.0	93 Np 237.0	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (260)

Useful Information

$$PV = nRT \quad R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}} = 8.314 \frac{\text{J}}{\text{mol}\cdot\text{K}}$$

$$P_{\text{solution}} = \text{solvent} P^{\circ}_{\text{solvent}} \quad \text{density of H}_2\text{O} = 1.00 \frac{\text{g}}{\text{cm}^3}$$

$$T = i k m \quad k_f(\text{H}_2\text{O}) = 1.86 \frac{^{\circ}\text{C}}{\text{m}} \quad k_b(\text{H}_2\text{O}) = 0.512 \frac{^{\circ}\text{C}}{\text{m}}$$

$$\ln \frac{k_1}{k_2} = \frac{E_a}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$$

$$\ln \frac{[A]_t}{[A]_0} = -kt \quad \frac{1}{[A]_t} - \frac{1}{[A]_0} = kt$$

$$K_p = K_c (RT)^n$$

$$x_{1,2} = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a} \quad \text{for } ax^2 + bx + c = 0$$

Temperature (°C)	Vapor Pressure(mmHg)	Temperature (°C)	Vapor Pressure(mmHg)
-5	3.2	50	92.5
0	4.6	55	118.0
5	6.52	60	149.4
10	9.20	65	187.5
15	12.8	70	233.7
20	17.5	75	289.1
25	23.8	80	355.1
30	31.8	85	433.6
35	42.1	90	525.8
40	55.3	95	633.9
45	71.9	100	760

Solubility Table

<u>Ion</u>	<u>Solubility</u>	<u>Exceptions</u>
NO ₃ ⁻	soluble	none
ClO ₄ ⁻	soluble	none
Cl ⁻	soluble	except Ag ⁺ , Hg ₂ ²⁺ , *Pb ²⁺
I ⁻	soluble	except Ag ⁺ , Hg ₂ ²⁺ , Pb ²⁺
SO ₄ ²⁻	soluble	except Ca ²⁺ , Ba ²⁺ , Sr ²⁺ , Hg ²⁺ , Pb ²⁺ , Ag ⁺
CO ₃ ²⁻	insoluble	except Group IA and NH ₄ ⁺
PO ₄ ³⁻	insoluble	except Group IA and NH ₄ ⁺
-OH	insoluble	except Group IA, *Ca ²⁺ , Ba ²⁺ , Sr ²⁺
S ²⁻	insoluble	except Group IA, IIA and NH ₄ ⁺
Na ⁺	soluble	none
NH ₄ ⁺	soluble	none
K ⁺	soluble	none

*slightly soluble

Metal	Half-Reaction Reaction
Lithium	Li $\text{Li}^+ + \text{e}^-$
Potassium	K $\text{K}^+ + \text{e}^-$
Barium	Ba $\text{Ba}^{2+} + 2\text{e}^-$
Calcium	Ca $\text{Ca}^{2+} + 2\text{e}^-$
Sodium	Na $\text{Na}^+ + \text{e}^-$
Magnesium	Mg $\text{Mg}^{2+} + 2\text{e}^-$
Aluminum	Al $\text{Al}^{3+} + 3\text{e}^-$
Manganese	Mn $\text{Mn}^{2+} + 2\text{e}^-$
Zinc	Zn $\text{Zn}^{2+} + 2\text{e}^-$
Chromium	Cr $\text{Cr}^{3+} + 3\text{e}^-$
Iron	Fe $\text{Fe}^{2+} + 2\text{e}^-$
Cobalt	Co $\text{Co}^{2+} + 2\text{e}^-$
Nickel	Ni $\text{Ni}^{2+} + 2\text{e}^-$
Tin	Sn $\text{Sn}^{2+} + 2\text{e}^-$
Lead	Pb $\text{Pb}^{2+} + 2\text{e}^-$
Hydrogen	H ₂ $2\text{H}^+ + 2\text{e}^-$
Copper	Cu $\text{Cu}^{2+} + 2\text{e}^-$
Silver	Ag $\text{Ag}^+ + \text{e}^-$
Mercury	Hg $\text{Hg}^{2+} + 2\text{e}^-$
Platinum	Pt $\text{Pt}^{2+} + 2\text{e}^-$
Gold	Au $\text{Au}^{3+} + 3\text{e}^-$