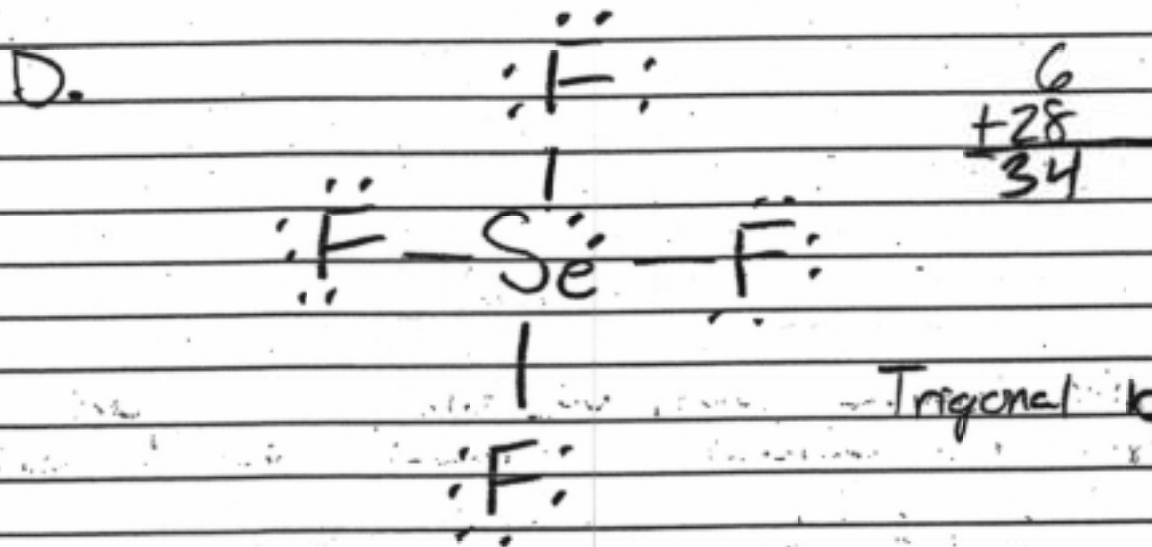


C. i) First ionization energy of Se is less than Br, because Se has one less proton pulling on its electrons. Br's e^- are being held w/a greater positive pull.

ii) Se has a great 1st ionization energy than Te, because Se experiences less shielding than Te. Te's e^- are held less tightly due to the shielding effect. This makes it easier for Te's e^- to be ionized.



~~This~~ This is a nonpolar molecule, because of the 4 Fs. These 4 atoms all have the same charge, so there is no polarity in the molecule.