

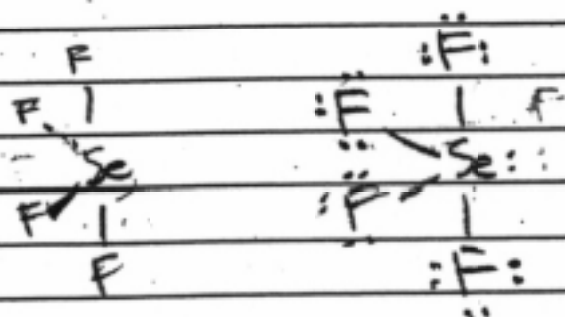
7ci continued
and harder to overcome.

7cii Se has a greater pull on its electrons than Te because Se has fewer layers of shielding electrons between the nucleus and the outermost subshell. Since there is less shielding on Se, its pull on its electrons is harder to overcome and it requires more energy for ionization.

7d. Se 6 valence electrons

F 7 valence electrons

$$6 + 4 \times 7 = 6 + 28 = 34 \text{ total electrons}$$



The molecule has a seesaw shaped structure and is therefore polar. Since the F atoms are arranged asymmetrically, the individual bond polarities do not cancel out and the molecule is polar.