

(ii) eq [H₂S]

Due to the 2:2 mole ratio the [H₂S] = [H₂]

so, [H₂S] = .173 M

c) $K_c = \frac{[H_2][S_2]}{[H_2S]^2}$

$K_c = \frac{[.173][.0275]}{[.175]^2}$

$K_c =$

d) partial pressures

$P_{S_2} = P - P_{H_2} - P_{H_2S}$

e) $K_p = \frac{[H_2S]}{[H_2][S_2]^{1/2}}$